

Triplet state dynamics in poly(2,5-pyridine diyl)

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Abstract

Pulse radiolysis was used to determine the triplet state energy (2.3 eV) of Poly(2,5-pyridine diyl) (PPY) film which was found to be coincident with the emission energy. Detailed time-resolved photoluminescence (PL) and pump-probe measurements have been applied to study the photoexcited state relaxation dynamics. In films, a very large spectral red-shift (~ 0.35 eV) for the PL occurred within the first 100 ps whereas no spectral red-shift was observed for the PPY in solution. This result shows clear evidence for the evolution of short-lived singlet emission ($S_1 \rightarrow S_0$) at ~ 450 nm to long-lived triplet emission ($T_1 \rightarrow S_0$) at ~ 520 nm for the PPY thin film. Streak camera measurement indicates the long live component has a decay time constant of several ns. The picosecond photo-induced triplet state absorption ($T_1 \rightarrow T_2$ transition) peaks at ~ 600 nm as measured by pump-probe which is consistent with both the radiolysis and cw photo-induced absorption measurements. A triplet lifetime of ~ 6 ns is measured which is again consistent with the streak camera measurement. These results lead us to believe that the long live component of the emission from PPY film is in fact phosphorescence. Furthermore, there is evidence that oxygen plays a very important role in the fast triplet radiative lifetime in PPY films. © 2001 Elsevier Science B.V. All rights reserved.

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Excited states and their relaxation processes in conjugated polymers have been of significant experimental and theoretical interest in the past decade. With the recent demonstration of polymer light emitting diodes (PLED) [1] and near commercialisation of these devices, a better understanding of the excited state properties of conjugated polymers became even more important. Luminescence in PLED is normally via recombination of singlet excitons. Spin statistics dictate that there is a 3:1 probability of triplet exciton compared to singlet exciton formation from electrical injected charge carrier recombination in PLED, thus a polymer in which efficient radiative triplet state recombination can occur will greatly enhance the PLED performance. Indeed a recent realisation of radiative triplet state recombination (i.e. phosphorescence) in organic light emitting device with a quantum efficient as high as 7% [2] compared to 3–4% for the singlet exciton emission PLED demonstrated the important role of the triplet state. Therefore, understanding the triplet state manifold and singlet–triplet intersystem crossing for the light emitting polymeric systems are important.

The existence of triplet states in conjugated polymers has been well established [3–6]. The photo-induced triplet state absorption ($T_1 \rightarrow T_2$ transition) measurements were the most frequent experimental technique to investigate these triplet states, however, direct triplet exciton emission (i.e. $T_1 \rightarrow S_0$) with microsecond lifetime has been observed in poly(3-hexylthiophene) [7] and ladder-type poly(paraphenylene) [8]. In this paper, we report the measurement of the singlet ground state (S_0) and first triplet level (T_1) of Poly(2,5-pyridine diyl) (PPY) using pulse radiolysis. The excited state relaxation dynamics were also investigated using femtosecond time-resolved luminescence and pump-probe techniques. These measurement results lead us to speculate that part of the emission from PPY film is in fact phosphorescence with a lifetime of ~ 6 ns.

The PPY is a good electron transport material as well as an efficient luminescent polymer, and it has been successfully used in PLED applications [9,10]. However, PPY is somewhat different from many other luminescent conjugated polymers, it has large Stokes shifts between the absorption and emission bands, higher photoluminescence (PL) efficiency in film compared to solution, and is stable in air. Fig. 1 shows the absorption and PL spectra for both solution and film. It is clearly show that the Stokes shift between the

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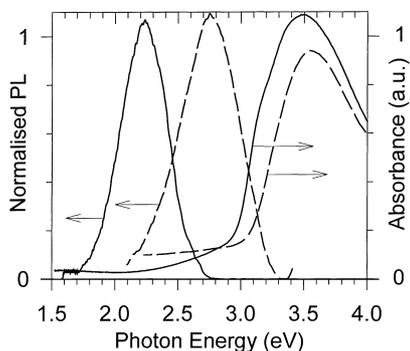


Fig. 1. Absorption and emission spectra of poly(2,5-pyridine diyl) film (solid line) and formic acid solution (dotted line) showing the large shifts in the relative positions of the features and the two states.

absorption and PL peaks is ~ 1 eV for the PPY film. This has been ascribed to either aggregate formation [11] or excimer formation [12]. In a recent study of the effects of protonation on the photophysics of PPY, we have been able to rule out both of these hypotheses [13] but were unable to unambiguously determine the cause of these large shifts. Further, and more importantly, the photoluminescence quantum yield (PLQY) of PPY behaves very differently from other polymers. The PLQY in the solid state (0.37) is more than double that in (formic acid) solution (0.17). Moreover, in the solid

state, the emission is stable in air, but decays rapidly in vacuum, slowly reaching the solution value, although the film absorbances do not alter [14]. This cannot be explained by any previous theory for conjugated polymers.

Pulse radiolysis is a very effective technique to determine the triplet energy and it has been recently applied to determine the S_0 - T_1 energy separation for soluble polymers [15]. By applying this pulse radiolysis and energy transfer technique with different energy acceptors, we have also measured the S_0 - T_1 energy separation for regioregular PPY and find it to be 2.4 eV in benzene solution. Thus, we estimate that with PPY in the solid state, this S_0 - T_1 energy difference is around 2.3 eV. This is just above the energy at which we observe the peak emission from PPY films. Therefore, there is reason to believe that emission in PPY films is possibly predominantly phosphorescence. Furthermore, theoretical calculations of the triplet manifold of PPY [16] yield results in excellent agreement with our radiolysis experiments. Details of the pulse radiolysis measurements of the triplet states in PPY can be found in the paper presented in this conference by Monkman et al. [17].

The above hypothesis is further supported by the femtosecond dynamics of the photoexcited states. Femtosecond time-resolved PL experiments were performed using a regeneratively amplified Ti:Sapphire laser in conjunction with frequency up-conversion; the time resolution of the

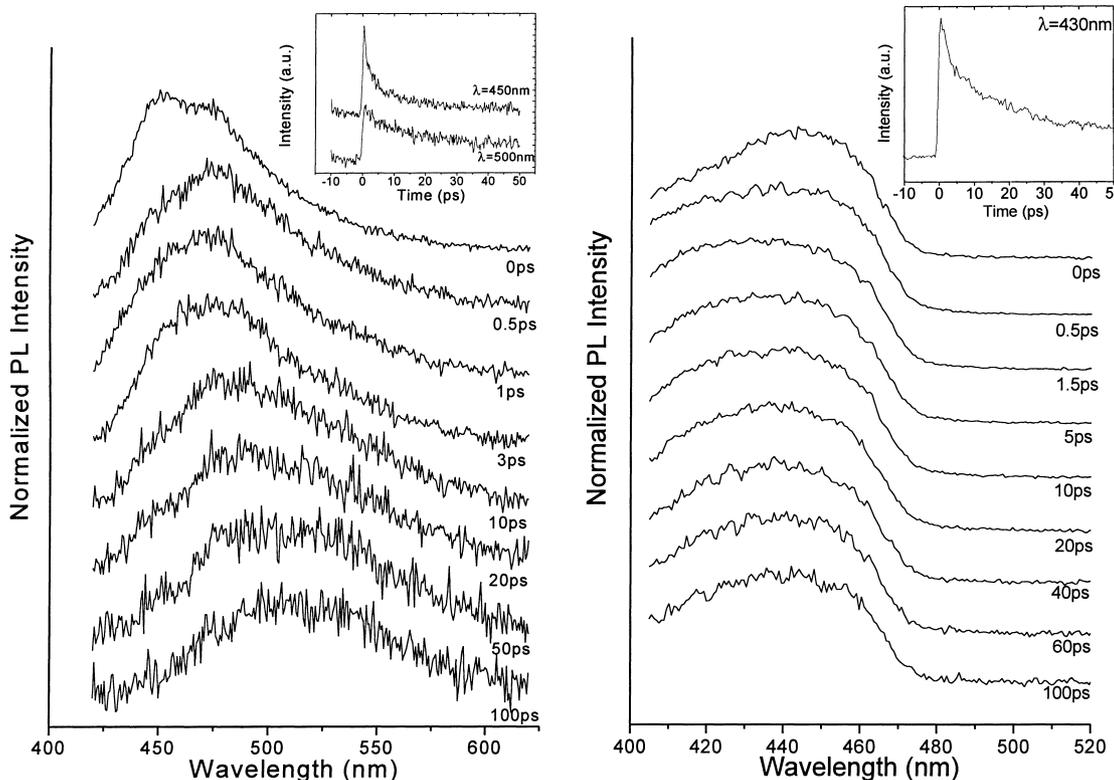


Fig. 2. Spectrally resolved femtosecond time resolved dynamics of the emission from poly(2,5-pyridine diyl) in the solid state. Initial very rapid decay of emission at 450 nm is observed with the build in of a long-lived component at ca. 525 nm. Inset shows the decay curves measured at both 450 and 500 nm. (b) In the solution state only the rapidly decay emission at 440 nm is observed, no long lived 525 nm emission is seen within our time measured time scale (up to 300 ps). Inset shows the decay curve recorded at 430 nm, unlike the thin film case the decay curves at other wavelengths remain the same.

system is ca. 300 fs and excitation energy 3.1 eV. Time resolved emission spectra from 0 to 100 ps delay of both film and solution are shown in Fig. 2. We observed a rapid and large red-shift of the film PL but no noticeable spectral red-shift for the solution case. In the case of films in air, we observed a very rapidly decay of the solution like emission peak at 450 nm, accompanied by a very long lived species peaking at ca. 525 nm which is similar to the time-integrated luminescence spectrum (note, films were used for many hours under continuous illumination in the open lab with very little sign of degradation). The figure insets show examples of the PL decay. For film, PL decay is multi-components and has strong wavelength dependence whereas the PL decay is essentially wavelength independent for the solution case. For PPY thin film at short wavelengths, i.e. 450 nm or less, the fast decay components (few ps to few 10 ps) dominated, however, a long decay component can still be clearly seen. Measuring emission at 525 nm, we observed mostly the long-lived species with a small contribution from the fast decaying component. We can clearly measure the long lived component by using a streak camera which can extend the time-resolved PL range to several nanoseconds. This measurement verified that the long-lived component has a lifetime of ~ 5 ns. Thus, we conclude that the very fast signal arises from the radiative decay of singlet states with the long-lived species being due to the phosphorescence. Fig. 3 shows the time-resolved emission peak position as a function of time. The red shift from short delay time emission at 450 nm to ca. 515 nm in 100 ps, a shift of some 0.35 eV is far too large to be due to spectral diffusion alone [18,19], especially as the PPY is a rigid rod polymer with a polydispersity of 1.7 [20]. Instead, it simply reflects the dominance of the emission band by singlet emission at short times and triplet emission at longer times.

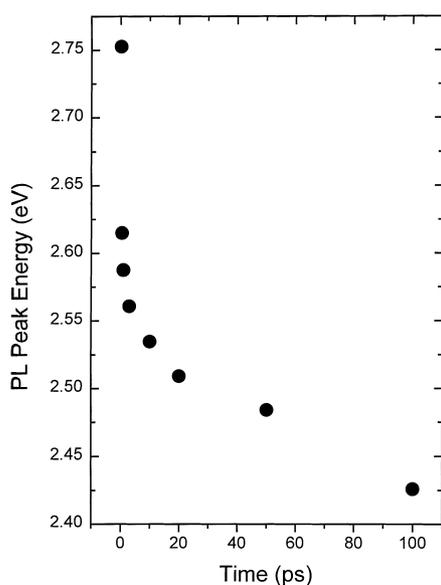


Fig. 3. The change in the emission peak energy as a function of time for the time-resolved PL spectra.

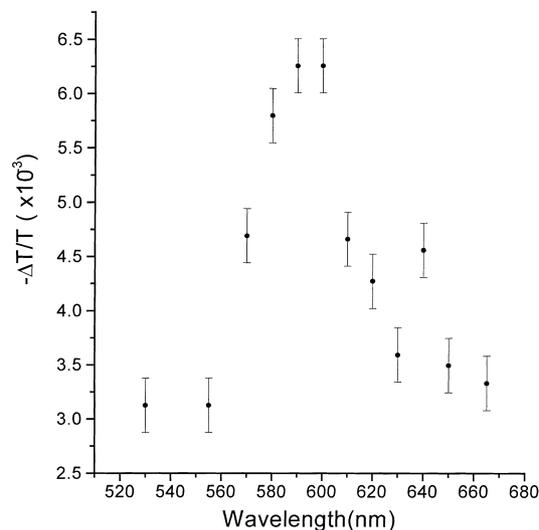


Fig. 4. The photo-induced absorption spectrum of PPY film at 100 ps time delay.

Thus, the very short lifetime of the singlet (τ in few ps to few 10 ps) reflects in part the usual case of spectral diffusion for this type of conjugated polymers and the competition between singlet emission and resonant ISC in PPY.

Fig. 4 shows the picosecond photo-induced absorption (PIA) of the PPY film at 100 ps time delay. The spectrum peak position and shape remain essentially unchanged at other time delays. It has a PIA peak at ~ 600 nm which is consistent with the T_1 - T_2 absorption peak observed in pulse radiolysis [17]. This result confirms that the picosecond PIA is due to triplet state absorption. Fig. 5 shows the PIA decay at wavelength of 610 nm, a single exponential can be fitted to the decay yielding a lifetime of ~ 6.5 ns. This triplet state lifetime is consistent with the phosphorescence decay time (~ 5 ns) measured by streak camera. Since a direct generation of triplet excitons by photoexcitation is forbidden, the triplets are normally created by singlet excitons via inter-

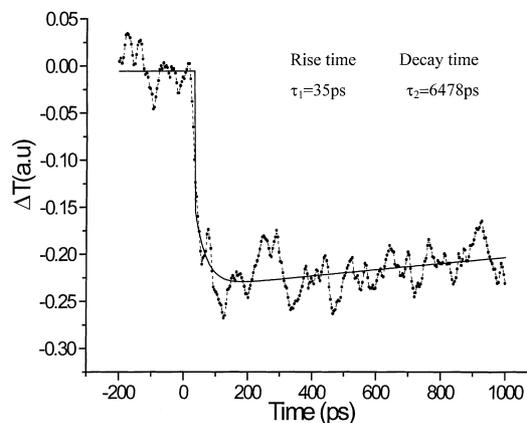


Fig. 5. The photo-induced absorption decay at $\lambda = 610$ nm. The solid line fitting curve using a exponential rise and decay function with time constant τ_1 and τ_2 , respectively.

system crossing (ISC) or singlet fission process [8]. The ISC process occurs in 10^{-10} – 10^{-9} s timescale [21], the PIA rise time shown in Fig. 5 is few tens of ps which seems to rule out this process. The singlet exciton fission process would be very fast but it requires the excitation energy to be twice that of the triplet energy [22–24], our present excitation energy at 3.1 eV again make this process unlikely. One possibility of fast triplet generation is by resonant ISC. It is well documented that upon photoexcitation both the singlet [25] and triplet [26] excited states of pyridine have distorted pyridyl ring structures, i.e. a prefulvenic structure [25]. That is the nitrogen atom bending up and out of the plane of the pyridyl ring. This distortion of the ring will enhance mixing of the n and π states, increasing the probability of ISC. Density Functional Theory calculations on the PPY system [16] show that the $^3(n-\pi^*)$ triplet state is degenerate in energy with the $^1(\pi-\pi^*)$, thus this can facilitate resonant (allowed) ISC.

This brings us to the question why triplet state emission appears in PPY film but not in solution, although triplet states are clearly seen in PPY solution by pulse radiolysis measurement. For an explanation, we need again turn to the parent monomer, it has been shown that O_2 complexes with the pyridine [27], and that the spin orbit coupling thus introduced perturbs the molecule such that radiative decay of the triplet ground state becomes allowed. Therefore, in polymer films, radiative triplet emission is caused by O_2 complexation. However, this complexation must occur in the excited state, i.e. when the chromophore is in the prefulvenic structure, since no direct optical absorption to the triplet state is seen in the polymer. In solution, the two processes will occur to reduce phosphorescence. Protonation does not greatly effect $(\pi-\pi^*)$ transitions whereas it does effect $(n-\pi^*)$ states, strongly shifting the $^3(n-\pi^*)$ state out of resonance with the $^1(\pi-\pi^*)$ state, such that ISC will be greatly reduced. Also the oxygen complexation site will be blocked on protonation so that the triplet emission is quenched, leaving mainly the singlet emission at 2.6 eV. The pulse radiolysis measurements show that in solution the triplets have a very long lifetime, again consistent with the switching off of their radiative decay channel. This also answers the perplexing question as to the instability of PPY emission in vacuum [14]. In air, the relatively triplet lifetime is short (order of 5–6 ns) as shown by our picosecond PIA and streak camera measurements, thus there is a much reduced probability of triplet energy transfer to O_2 to form singlet oxygen, giving rise to the high stability to photooxidation of the polymer.

In conclusion, pulse radiolysis measurements have determined the triplet state energy of PPY film to be coincident with the emission energy. Detailed time-resolved PL and pump-probe measurements have been applied to study the

photoexcited state relaxation dynamics. A very large spectral red-shift (~ 0.4 eV) for the PL occurred within the first 100 ps for the film case whereas no spectral red-shift was observed for the PPY in solution. This result shows clear evidence for the evolution of short-lived singlet emission ($S_1 \rightarrow S_0$) at ~ 450 nm to long-lived triplet emission ($T_1 \rightarrow S_0$) at ~ 520 nm for the PPY thin film. Furthermore, streak camera measurement indicates the luminescence has a very long live component with a decay constant of several ns. The photo-induced triplet state absorption ($T_1 \rightarrow T_2$ transition) peaks at ~ 600 nm measured by pump-probe are consistent with the radiolysis results. A triplet lifetime of ~ 6 ns is measured which is again consistent with the streak camera measurement. These results lead us to believe that the long live component of the emission from PPY film is in fact phosphorescence.

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