

Synthesis, crystal structure and magnetic properties of an alternating manganese chain

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Received 2 February 2006; received in revised form 15 March 2006; accepted 17 March 2006

Available online 28 March 2006

Abstract

A new 1D complex has been prepared and characterized. X-ray single crystal structure confirms that the Mn(II) ions assemble in alternating chains with Mn–Mn distances of 3.8432(13) and 4.4428(14) Å. A 3D network of hydrogen bonds links the chains together. The temperature dependence of the magnetic susceptibility reveals that this compound undergoes a magnetic transition and exhibits an antiferromagnetic interaction in the low-temperature phase with two alternating exchange interactions of $-2.32(1)$ and $-5.55(1) \text{ cm}^{-1}$.

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Keywords: Mn(II) complexes; Dimethylglycine; X-ray structure; Antiferromagnetic interaction

1. Introduction

The study of paramagnetic manganese carboxylate clusters was given a boost since the discovery that some molecules can behave as nanoscale magnets [1]. At low enough temperatures, those molecules display a magnetization versus applied field hysteresis loops associated with a slow magnetization relaxation. Their properties derive from the combination of a large ground-spin (S) value and a significant magnetoanisotropy of the Ising (easy-axis) type. Many applications of these single-molecule magnets (SMMs) are devised, such as in high-density information storage, where each bit of information would be stored on a single molecule, and in quantum computing, where each molecule would function as a quantum bit. Of the SMMs known to date, the archetypal Mn_{12} family constitutes the most famous and studied example with $[\text{Mn}_{12}\text{O}_{12}(\text{O}_2\text{CR})_{16}(\text{H}_2\text{O})_4]$ exhibiting SMM behaviour at the highest temperature (ca. 6 K) [2].

SMM behaviour has been found not only in cluster type compounds but also in 1D chains. Tetranuclear manganese units $(\text{Mn}_4(\text{hmp})_6)$ bridged by chloride ions [3] or chains of ferromagnetic coupled Mn(III)–Ni(II)–Mn(III) units [4], have shown slow magnetization relaxation.

Carboxylate-bridged complexes have shown to be versatile, both in the bridging coordination modes of the carboxylate and in the resulting magnetic properties. A loose correlation has even been found between the type and number of bridging mode and the magnetic interactions. Durot et al. [5] report that in Mn(II) complexes the magnitude of the antiferromagnetic coupling mediated through carboxylato bridges ($|J|$) increases with the number of bridges, with the *syn-syn* triatomic mode ($-\text{OCO}-$) being the more efficient pathway. $|J|$ values for halogenato Mn(II) chains are usually higher [5].

In this paper we describe the synthesis of a dimethylglycine and chloride-bridged Mn(II) compound with an antiferromagnetic alignment of the metallic spins, below 23 K. We have essayed the synthesis of compounds with dimethylglycine and manganese driven by the success in obtaining low-dimensional compounds with trimethylglycine and manganese, namely chains [6,7] and SMMs clusters [8].

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2. Experimental

2.1. Synthesis of catena triaquatrichloro- μ_2 -(bis(*N*-dimethylglycine-*O*,*O'*)) dimanganese(II) chloride monohydrated

0.88 g of dimethylglycine were added to 1.76 g of hydrated manganese chloride in a water solution (2 mL). After a few months at room temperature and pressure small single crystals were grown, with a pale pink colour.

2.2. X-ray structure determination

A single crystal with dimensions $0.24 \times 0.18 \times 0.07$ mm³ was used for structural determination at 293 K on an Enraf–Nonius diffractometer with a point detector and graphite-monochromated MoK α radiation. All computations were carried out on a PC-586 computer running Linux. Number of measured, unique, and observed reflections [$I > 2\sigma(I)$] are 3782, 3501 and 2710 ($R_{\text{int}} = 2.52$). The structure was solved by direct methods using SHELXS97 [9] and refined on F^2 by full-matrix least-squares methods using SHELXL97 [9]. All the non-hydrogen atoms were refined anisotropically. Hydrogen atoms were constrained to ride on the respective parent atoms with an isotropic displacement parameter multiple of that of the parent atom using SHELXL97 defaults. Crystal data are summarized in Table 1. Selected bond lengths and bond angles are listed in Table 2. H-bond geometry is listed in Table 3.

2.3. Magnetic measurements

Magnetic data on crushed single crystals were collected using a Quantum-Design MQMS superconducting

Table 1
Crystallographic data

Formula	C8 H26 Cl4 Mn2 N2 O18
M	529.99
Crystal System	Monoclinic
Space Group	$P2_1/c$
a (Å)	7.5801(19)
b (Å)	16.324(5)
c (Å)	16.857(4)
β (deg)	106.21(3)
V (Å ³)	2002.9(10)
Z	4
D_c (g cm ⁻³)	1.758
$F(000)$	1080
Goodness-of-fit on F^2	1.089
R_1 , wR_2 [$I > 2\sigma(I)$] ^a	0.0233, 0.0639
(All data)	0.0491, 0.0576
$(\Delta\rho)_{\text{min}}$, $(\Delta\rho)_{\text{max}}$ (e ⁻ Å ⁻³)	-0.317, 0.410

^a $R_1 = \Sigma ||F_0| - |F_c|| / \Sigma |F_0|$, $w = 1 / [\sigma^2(F_0^2) + (0.0278P)^2 + 1.4570P]$ where $P = (F_0^2 + 2F_c^2) / 3$.

Table 2
Selected geometrical parameters

Bond distances (Å)	
O1–C1	1.258(3)
N1–C4	1.485(4)
N1–C2	1.485(3)
N1–C3	1.493(4)
O2–C1	1.240(3)
O3–C5	1.241(3)
O4–C5	1.249(3)
N2–C7	1.477(3)
N2–C6	1.478(3)
N2–C8	1.484(3)
C1–C2	1.515(3)
C5–C6	1.515(3)
Bond angles (deg)	
Mn1–Cl1–Mn2	99.21(3)
Mn1–Cl2–Mn2	119.57(3)
O2–C1–O1	126.8(2)
O3–C5–O4	127.2(2)
Torsion angles (deg)	
C8–N2–C6–C5	-165.0(2)
C7–N2–C6–C5	71.4(3)
C3–N1–C2–C1	77.2(3)
C4–N1–C2–C1	-158.4(3)
O2–C1–C2–N1	10.9(3)
O1–C1–C2–N1	-171.0(2)
N2–C6–C5–O3	27.6(3)
N2–C6–C5–O4	-154.0(2)

Table 3
H-bond parameters

Atoms	D–H	H...A	D–H...A	Angles
O6–H61...O1 ⁱ	0.84(4)	1.93(4)	2.769(3)	176(3)
O6–H62...O4	0.79(4)	2.47(4)	2.872(3)	113(3)
O6–H62...Cl4 ⁱⁱ	0.79(4)	2.62(4)	3.344(2)	153(3)
O7–H71...Cl4	0.82(4)	2.41(4)	3.224(2)	178(3)
O7–H72...O4 ⁱⁱⁱ	0.74(4)	2.05(4)	2.774(3)	171(4)
O8–H81...O6 ^{iv}	0.88(5)	2.26(5)	2.943(4)	135(4)
O8–H82...Cl2 ^v	0.89(5)	2.43(5)	3.278(3)	159(4)
O5–H51...Cl3 ^{vi}	0.80(4)	2.36(4)	3.144(2)	167(4)
O5–H52...Cl4 ^{vi}	0.86(4)	2.27(4)	3.119(2)	169(3)
N1–H1...O8 ^{vii}	0.86(3)	1.95(3)	2.748(4)	155(3)
N2–H2...Cl4	0.88(3)	2.39(3)	3.181(2)	150(2)

i: $x+1, y, z$; ii: $-x, -y, -z+1$; iii: $x-1, y, z$; iv: $x-1, -y+3/2, z-1/2$; v: $-x, -y+1, -z$; vi: $x+1, -y+3/2, z+3/2$; vii: $1+x, 1/2-y, 1/2+z$.

quantum interference device (SQUID) magnetometer. Susceptibility curves were measured over the temperature range 2–250 K, after field cooling (100 Oe) and zero-field cooling. Magnetization data was recorded at several temperatures using magnetic fields from 0 to 5.5 T.

3. Results and discussion

3.1. Crystal structure

The metallic atoms are assembled in infinite chains of dimers running along the *a*-axis. Each dimeric unit contains two crystallographically independent Mn(II) ions bridged by the carboxylic groups of two organic moieties and one chloride ion (Fig. 1). A distinct chloride ion links the dimers into chains (Fig. 2). Both metallic ions have approximated octahedral geometries.

Mn1 is surrounded by two carboxylic oxygen atoms, 1 water oxygen and 3 chloride ions, with the Cl⁻ ions being

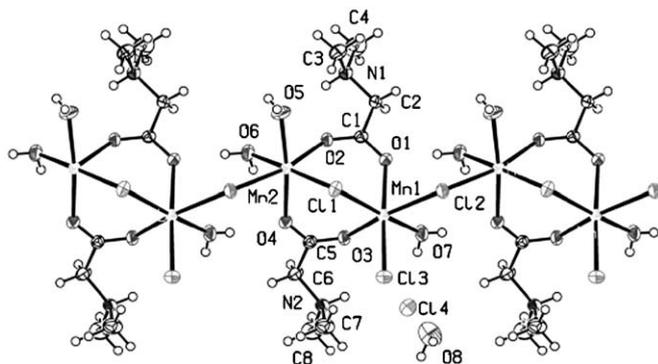


Fig. 1. Ortep diagram of the title compound. Ellipsoids are drawn at the 50% probability level.

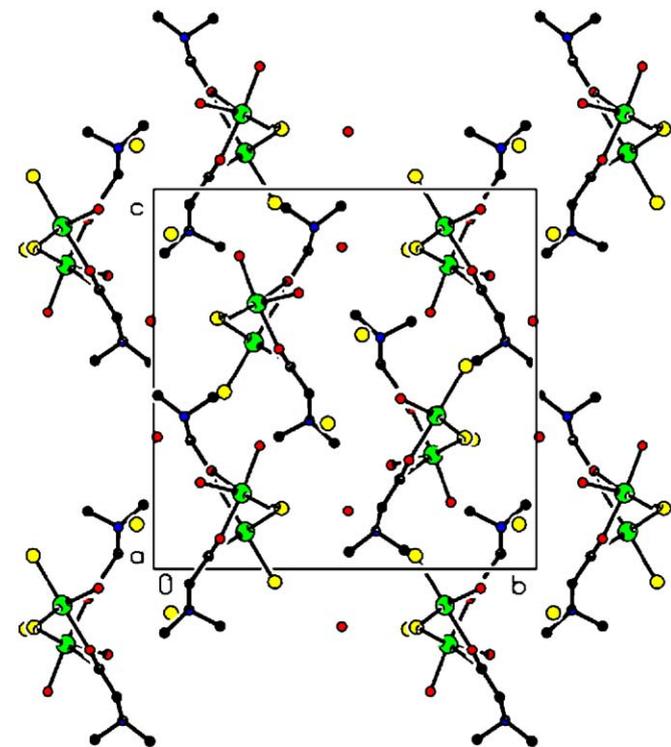


Fig. 2. Two of the infinite chains that run along the *a*-axis. Hydrogen atoms were omitted for clarity.

mutually *cis*, Cl–Mn1–Cl angles ranging from 91.79(4)° to 97.07(3)°.

Mn2 is coordinated by 2 carboxylic oxygen atoms, 2 water oxygen atoms, and 2 chloride ions, with Cl1–Mn2–Cl2 angle of 96.29(3)°.

The inner-dimer bridging angle, Mn1–Cl1–Mn2, is 99.22(4)°. Along the chain, Mn1 has two nearest Mn2 neighbours at distances 3.8432(13) and 4.4428(14) Å. The outer-dimer bridging angle, Mn1–Cl2–Mn2^{*i*}, is 119.57(3)° (*i*: $-1+x, y, z$).

The organic moieties (substituted glycine molecules) are present in the neutral, dipolar, zwitterionic form. Inspection of the C–O distances in the carboxylic groups show that these groups are deprotonated. Carboxylic groups are rotated about 10°[26°] around C1–C2 and C5–C6, respectively. Chains are interconnected through hydrogen bridges in a complex 3D network of H-bonds (Fig. 3). Each chain establishes hydrogen bridges with 5 other chains. Cl4 ion accepts protons donated by 3 neighbouring chains. The isolated water molecule shares its hydrogen atoms with another water oxygen (O6) and a chloride ion (Cl2) of the near chains and accepts another from the amino group of one of the substituted aminoacids, N1.

The formation of chains is also seen in a similar compound, *catena diaqua-dichloro-μ₂-(methylglycine-O,O')* manganese (II), previously reported by us [10]. In this case a similar synthetic procedure led to the formation of uniform *zig-zag* chains with the metal atoms solely bridged by the organic moieties.

3.2. Magnetic properties

The temperature dependence of the magnetic susceptibility of *catena triaquatrichloro-μ₂-(bis(N-dimethylglycine-O,O'))*

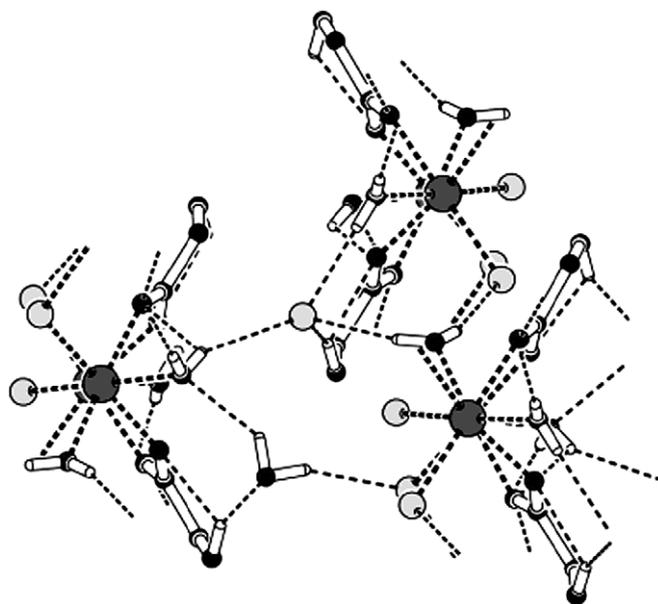


Fig. 3. H-bonding network. H-bonds are drawn in light-dashed lines.

dimanganese(II) chloride monohydrated was measured under an applied field of 100 Oe in the temperature range 2–250 K, both in the field-cooled and zero field-cooled regime, with very similar results.

The plot of the molar susceptibility versus T is shown in Fig. 4. The overall magnetic behaviour corresponds to an antiferromagnetic-coupled system. The increasing of the susceptibility at very low temperatures is attributed to the presence of impurities. Fig. 5 plots χ^{-1} versus T . Upon cooling down the sample, χ^{-1} linearly decreases until 50 K, as expected for a paramagnetic behaviour. The high-temperature range ($T > 50$ K) can be well fitted by the equation $\chi = (C/T - \theta) + \text{TIP}$ with the fitting parameters being $C = 9.32(5) \text{ cm}^3 \text{ K mol}^{-1}$, $\theta = -52.5(4) \text{ K}$ and $\text{TIP} = 6(2) \times 10^{-4} \text{ cm}^3 \text{ mol}^{-1}$.

A least-squares fitting of the Cortés expansion [11]:

$$\chi = \frac{Ng^2\mu_B^2}{3kT} \left(\frac{1 + u_1 + u_2 + u_1u_2}{1 - u_1u_2} \right)$$

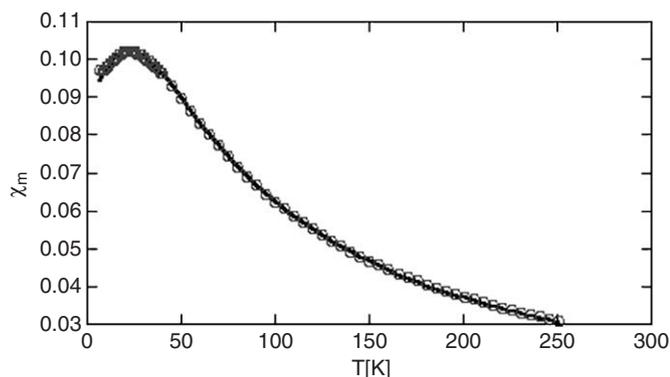


Fig. 4. Open circles represent the magnetic susceptibility experimental values. The solid black line represent the theoretical Cortés fit as explained in the text.

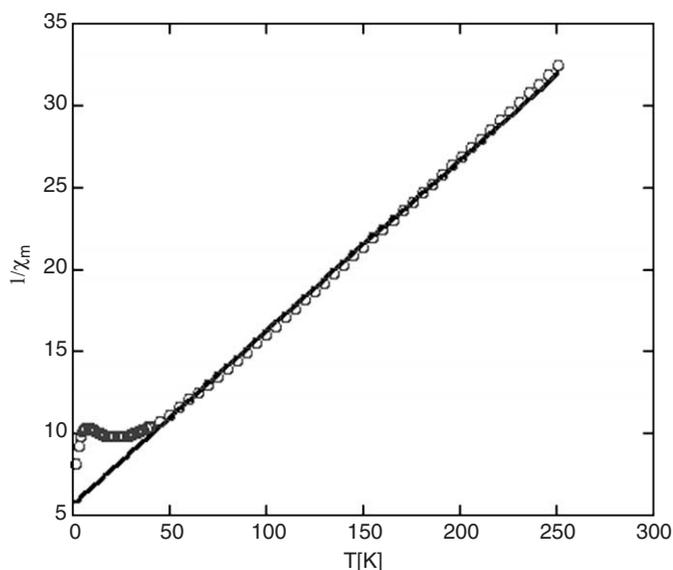


Fig. 5. Thermal variation of the inverted magnetic susceptibility ($\text{cm}^{-3} \text{ mol}$). The solid line is a fit of the paramagnetic region using the Curie-Weiss model.

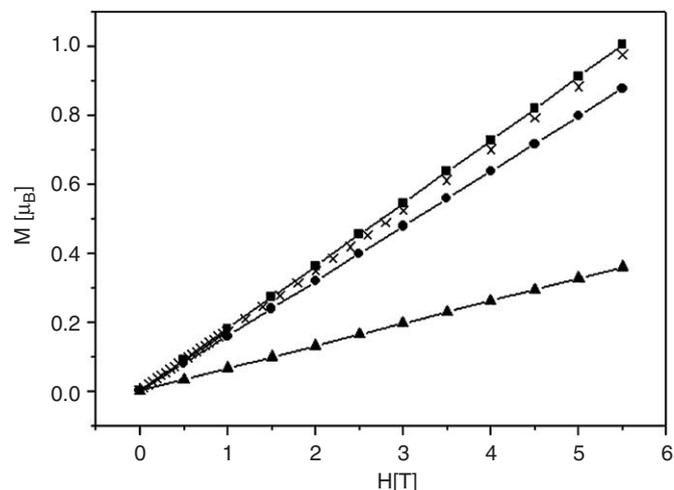


Fig. 6. Magnetization (μ_B) curves versus applied magnetic field (T) at 10 K (x), 25 K (■), 50 K (●) and 200 K (▲).

with $(\coth(J_p/kT) - kT/J_p) = \mu_p$, returned as best parameters $J_1 = -2.32(1) \text{ cm}^{-1}$, $J_2 = -5.55(1) \text{ cm}^{-1}$ and $g = 2.000(2)$. The fitting was performed in the 7–250 K region, and fits the data perfectly. The theoretical curve is represented by a solid line in Fig. 4.

Fig. 6 shows the magnetization curves measured at 10, 25, 50 and 200 K with the applied magnetic field ranging from 0 to 5.5 T. The linear dependence of the magnetization of the sample with the applied field is characteristic of an antiferromagnetic interaction.

4. Supplementary materials

Crystallographic data for structural analysis have been deposited with the Cambridge Crystallographic Data Center, CCDC No. 297113. Copies of this information may be obtained free of charge on application to CCDC, 12 Union Road, Cambridge CB2 1EZ, UK (fax: +44 1223 336 033; e-mail: deposit@ccdc.cam.ac.uk or <http://www.ccdc.cam.ac.uk>).

Acknowledgments

This work was supported by Fundação para a Ciência e a Tecnologia under project POCI/FIS/57876/2004. The authors are indebted to Prof. Margarida Godinho for the use of the SQUID magnetometer.

Appendix A. Supplementary materials

Supplementary data associated with this article can be found in the online version at [doi:10.1016/j.jssc.2006.03.024](https://doi.org/10.1016/j.jssc.2006.03.024).

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